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9.1 Oxidation and reduction - IB Alchemy

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Oxidation - The loss of electrons from a substance. This may happen through the gain of oxygen or loss of hydrogen. Reduction - The gain of electrons. This may happen though the loss of oxygen, or gain of hydrogen, or decrease in oxidation number. 9.1.2 - Deduce the oxidation number of an element in a compound IB Chemistry - Unit 9: Oxidation and Reduction (Redox ... 9.1 - Introduction to Oxidation and Reduction | A* Chemistry IB Chemistry Topic 9 Redox processes Topic 9.1 Oxidation and reduction SL Covers oxidation numbers, redox half equations, and titrations. Full set of resourc...

9.5.2 State that oxidation occurs at the positive electrode (anode) and reduction occurs at the negative electrode (cathode). Electrons flow from anode to cathode. Anode - positive charge, attracts negative ions. Oxidation happens at the anode because substances must lose electrons.

9.1 Deduce simple oxidation and reduction half-equations given the redox reaction [SL IB Chemistry] ... [SL IB Chemistry] - Duration: 2:54. Richard Thornley 34,646 views. 2:54.

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Syllabus ref: 9.1. Oxidation is defined as loss of electrons, reduction is defined as gain of electrons. Reactions that involve electron transfer can be described in terms of reduction and oxidation. Reduction and oxidation reactions; Oxidation; Oxidising agents; Reduction; Reducing agents; Redox reactions

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IB Chemistry Review: Topic 9: Oxidation and Reduction

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9.1 Deduce simple oxidation and reduction half-equations given the redox reaction [SL IB Chemistry]

9.1 - Introduction to oxidation and reduction . Oxidation is the loss of electrons, reduction is the gain of electrons. Oxidation half equation... Mg Mg 2+ + 2e-. Reduction half equation... O + 2e- O 2-. Calculation of oxidation number. There are a few rules to remember. The oxidation number of an element is zero.

IB Chemistry standard level revision notes: Reduction and ...

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