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Formation of Compounds

Chapter 4: Unit 3. Formation of Ionic Compounds. 3. Formation of Ionic Compounds We will discuss the formation NaCl ionic compounds. Ionic compounds are formed between a metal and a non-metal. Sodium for example is located under Group I. Therefore following octet rule, ...

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saw how the nature of bonding in ionic compounds explains these properties. Let us now study the properties of some carbon compounds. Melting and boiling points of some carbon compounds are given in Table 4.1. Most carbon compounds are poor conductors of electricity as we have seen in Chapter 2. From the data on the boiling and melting points ...

CHAPTER4 Carbon and its Compounds

Thus, the term molecular compound is used to describe elements that are covalently bonded and to distinguish the compounds from ionic compounds. Some pure elements exist as covalent molecules. Hydrogen, nitrogen, oxygen, and the halogens occur naturally as the diatomic ("two atoms") molecules H₂, N₂, O₂, F₂, Cl₂, Br₂, and I₂ (part (a) in Figure 4.1).

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